Rational co-catalyst design for amide hydrogenolysis based on DFT calculations.

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The reduction of electron rich carboxylic acid derivatives, such as amides, is still a difficult reaction. In a recently reported method, amide hydrogenolysis was enhanced by the combination of a secondary amide co-catalyst with (^{iPr}PNP)Fe(H)(CO) ($^{iPr}PNP = N[CH_2CH_2(P^iPr_2)]_2$).[1] The mechanism for this reaction was studied in our group by using DFT methods.[2] The M06 functional was selected based on a method benchmark using CCSD(T). The mechanistic study showed that the secondary amide co-catalyst acts as a proton shuttle in the rate limiting step (ΔG^{\ddagger}_{HT}) but poisons the catalyst forming an adduct (ΔG_{add}) (Figure 1). In order to improve this reaction, ΔG_{add} and ΔG^{\ddagger}_{HT} were calculated using as co-catalysts the molecules shown in Figure 1 and then introduced in a microkinetic model to predict the expected amide conversion.[3] The most promising candidates were tested in the lab and the experimental results were in good agreement with the computational predictions. Co-catalyst 1 yielded the highest conversion, owing to an optimal balance between ΔG_{add} and ΔG^{\ddagger}_{HT} .

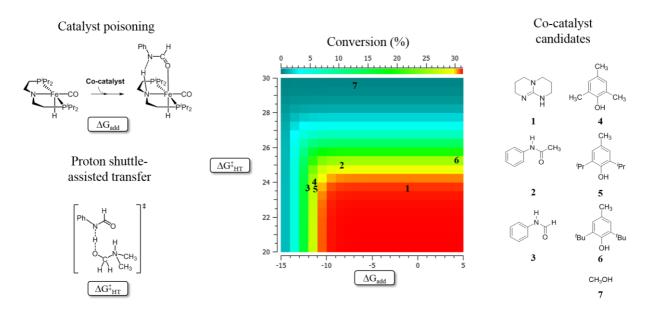


Figure 1: Key reactions in the amide hydrogenolysis co-catalyst design (left); predicted conversion vs ΔG_{add} and ΔG^{\dagger}_{HT} (middle); co-catalysts tested (right).

References

1. U. Jayarathne, Y. Zhang, N. Hazari and W.H. Bernskoetter, *Organometallics* (2017), **36**, 409

2. L. Artús Suàrez, W.H. Bernskoetter, N. Hazari and A. Nova, ACS Catal. 8 (2018), 8751

3. L. Artús Suàrez, U. Jayarathne, D. Balcells, W.H. Bernskoetter, N. Hazari, M. Jaraiz and A. Nova, *ACS Catal*, Submitted for publication.